

Chapter 11 The Mole Answer Key

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The calculated answer (0.5639 mol) is indeed less than 1 mol, so we have probably not made a major error in the calculations. C To calculate the number of molecules in the sample, we multiply the number of moles by Avogadro's number: moleculesof ethylene glycol = 0.5639 mol(6.022 × 10²³ 1 mol) = 3.396 × 10²³ molecules.**

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Chapter 11: The Mole

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dividing the mass of the element in one mole of the compound by the mass of one mole of the compound and multiplying by 100. if there is more than one of the element in in the compound like ch₄ has 4 hydrogen, then you have the multiply the mass of hydrogen by 4 and then divide it by the mass of the ch₄ and then multiply by 100.

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